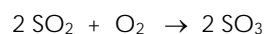


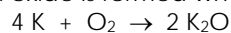
# Stoichiometry Practice

Indicate the state: *gas, liquid, solid, aqueous* of each substance.

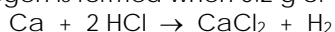
- 1) What mass of sulphur trioxide is formed from 96 g of sulphur dioxide?



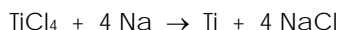
- 2) What mass of potassium oxide is formed when 9.75 g of potassium is burned in oxygen?



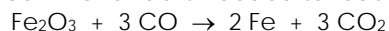
- 3) What mass of hydrogen is formed when 0.2 g of calcium reacts with hydrochloric acid?



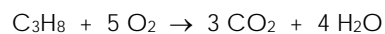
- 4) What mass of sodium is needed to reduce 1 kg of titanium chloride?



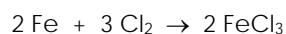
- 5) What mass of carbon monoxide is needed to reduce 1 kg of iron oxide to iron?



- 6) What mass of oxygen is needed to burn 110 g of propane ( $\text{C}_3\text{H}_8$ )?



- 7) What mass of iron reacts with 14.2 g of chlorine?



- 8) 4.17 g of hydrated barium bromide crystals ( $\text{BaBr}_2 \cdot n\text{H}_2\text{O}$ ) gave 3.72 g of anhydrous barium bromide on heating to constant mass. Work out the relative molecular mass ( $M_r$ ) of the hydrated barium bromide and the value of  $n$ .

